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A sample of calcium carbonate is heated. $\text{CaCO}_3(\text{s}) \rightarrow \text{CaO}(\text{s}) + \text{CO}_2(\text{g})$ Sulfur dioxide gas is bubbled through water. $\text{SO}_2(\text{g}) + \text{H}_2\text{O}(\text{l}) \rightarrow \text{H}_2\text{SO}_3(\text{aq})$ Solid potassium oxide is added to a container of carbon dioxide gas. $\text{K}_2\text{O}(\text{s}) + \text{CO}_2(\text{g}) \rightarrow \text{K}_2\text{CO}_3(\text{s})$ Liquid hydrogen peroxide is warmed. $2\text{H}_2\text{O}_2(\text{l}) \rightarrow 2\text{H}_2\text{O}(\text{l}) + \text{O}_2(\text{g})$ Solid lithium oxide is added to water.

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The Ultimate Chemical Equations Handbook Determine the oxidation number of the underlined element: KMnO_4 . Since K is an alkali metal, its charge must be $1+$. Oxygen is $2-$ but there are four of them, therefore, 4 times $2-$ equals $8-$. If $1+$ and $8-$ are added together, we get $7-$. In order for the compound to be neutral, the Mn must be $7+$.

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